***Proton***

Electron

* It’s a particle located in the \_\_\_\_\_\_\_\_\_\_\_\_\_ of an atom.

Neutron

* It has a charge of \_\_\_\_\_\_\_\_\_\_\_ \_\_\_ and a mass of \_\_\_\_\_

amu (atomic mass units).

* \_\_\_\_\_\_\_\_\_\_\_\_ are almost identical in size to neutrons.

The protons determine the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

Proton

* The \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of protons will tell you what element

it is.

* An atom having 6 protons will always be Carbon, C.
* The atomic number is the same as the number of protons.

***Electron***

* + It’s a particle located \_\_\_\_\_\_\_\_\_\_\_\_\_ the nucleus of an atom.
	+ It has a charge of \_\_\_\_ and a mass of \_\_\_ amu (atomic mass units).
	+ Electron are very small, they are \_\_\_\_\_\_\_\_\_\_\_ the size of a proton.
* In a neutral atom, the number of \_\_\_\_\_\_\_\_\_\_\_\_\_ is the same as the atomic number.
* If an atom becomes an \_\_\_\_\_\_\_, the number of electrons changes creating a charged atom.

***Neutron***

* + It’s a particle located \_\_\_\_\_ the nucleus of an atom.
	+ It has a charge of \_\_\_\_ and a mass of \_\_\_\_ amu (atomic mass units).
	+ Neutrons are almost identical in size to \_\_\_\_\_\_\_\_\_\_\_\_\_.
* Number of neutrons depends on the **\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**.

***Mass Number***

******

* Mass number is the number of \_\_\_\_\_\_\_\_\_\_\_\_ plus \_\_\_\_\_\_\_\_\_\_\_\_\_.
* Mass number is found by adding protons and neutrons.
	+ 3 protons + 4 neutrons = 7 (Lithium)
* Or can be found by rounding atomic mass to the nearest \_\_\_\_\_\_\_\_\_\_\_\_\_\_ number.
	+ Iodine atomic mass is 126.9044 so its mass number is 127.
* Mass number can be used to calculate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ in an atom.
	+ Iodine mass # 127 – Iodine atomic # 53 = 74 neutrons.

***Calculating Particles in an Atom***





