REVIEW

4

SECTION 4.2

Ionic and Covalent Bonding

1.	Explain why solid table salt is a poor conductor of electricity, while salt water is a good conductor of electricity.
2.	Explain why table salt does not melt easily.
3.	Contrast ionic and covalent bonds.
4.	Explain why a triple bond between two nitrogen atoms is stronger than a double bond between two oxygen atoms.
5.	Explain how it is possible for a compound to have both ionic and covalent bonds.
6.	Predict whether a gold ring would be a good conductor of electricity. What kind of bonds does gold have? How do these bonds explain gold's properties?